Solubility of Some Ionic Compounds in Water	
Always Soluble Alkali metals (Li^+, Na^+, K^+, Rb^+, Cs^+), NH_4^+ Nitrate = NO_3^-, Chlorate = ClO_3^-, Perchlorate = ClO_4^-, Acetate = $C_2H_3O_2^-$	mnemonics
Generally soluble Cl^-, Br^-, I^- Soluble except: Ag^+, Pb^{2+}, Hg_2^{2+} F^- Soluble except: Ca^{2+}, Sr^{2+}, Ba^{2+}, Pb^{2+}, Mg^{2+}	AP/H CBS-PM
SO_4^{2-} Soluble except: Ca^{2+}, Sr^{2+}, Ba^{2+}, Pb^{2+}	CBS-PBS
Generally Insoluble O^{2-}, OH^- Insoluble except: Alkali metals and NH_4^+ Ca^{2+}, Sr^{2+}, Ba^{2+} somewhat soluble CO_3^{2-}, PO_4^{3-}, S^{2-}, SO_3^{2-}, $C_2O_4^{2-}$, CrO_4^{2-} Insoluble except: Alkali metals and NH_4^+	CBS

Not Soluble = forms precipitate **Soluble** = dissolves in water (aqueous)